

SFSU Chem 100  
M. Corcoran  
Key Worksheet

Name \_\_\_\_\_

Chemical Reactions and Stoichiometry

1. Given the equation  $3A + B \rightarrow C + D$ , you react 1 mole of A with 3 moles of B. True or false: B is the limiting reactant because you have fewer moles of B than A.

Answer: False

$$\frac{1 \text{ mole A}}{3 \text{ mole A}} = \frac{1 \text{ mole B}}{3 \text{ mole B}} = 0.3 \text{ moles B needed. Have 3 moles B}$$

so A is limiting

2.  $\text{NaHCO}_3$  is the active ingredient in baking soda. How many grams of oxygen are in 0.35 g of  $\text{NaHCO}_3$ ?

- a) 0.012 g
- b) 0.022 g
- c) 0.067 g
- d) 0.20 g

ANS: d)

$$0.35 \text{ g NaHCO}_3 \left( \frac{1 \text{ mole NaHCO}_3}{84.007 \text{ g NaHCO}_3} \right) \left( \frac{3 \text{ moles O}}{1 \text{ mole NaHCO}_3} \right) \left( \frac{16.00 \text{ g}}{1 \text{ mole O}} \right) =$$

$$= 0.199 = \boxed{0.20 \text{ g}}$$

3. Suppose the reaction  $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$  is carried out starting with 103 g of  $\text{Ca}_3(\text{PO}_4)_2$  and 75.0 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

- a) 74.9 g
- b) 50.0 g
- c) 112 g
- d) 32.5 g

ANS: b)

$$103 \text{ g Ca}_3(\text{PO}_4)_2 \left( \frac{1 \text{ mole}}{310.1875 \text{ g Ca}_3(\text{PO}_4)_2} \right) = 0.332 \text{ moles Ca}_3(\text{PO}_4)_2$$

$$75.0 \text{ g H}_2\text{SO}_4 \left( \frac{1 \text{ mole H}_2\text{SO}_4}{98.082 \text{ g H}_2\text{SO}_4} \right) = 0.765 \text{ moles H}_2\text{SO}_4 \text{ L.P.}$$

$\text{Ca}_3(\text{PO}_4)_2$  needed:  $0.765 \text{ moles H}_2\text{SO}_4 \left( \frac{1 \text{ mole Ca}_3(\text{PO}_4)_2}{3 \text{ moles H}_2\text{SO}_4} \right) = 0.255 \text{ moles}$  Have excess

$$0.765 \text{ moles H}_2\text{SO}_4 \left( \frac{2 \text{ moles H}_3\text{PO}_4}{3 \text{ moles H}_2\text{SO}_4} \right) \left( \frac{97.994 \text{ g H}_3\text{PO}_4}{1 \text{ mole}} \right) = 49.97 \text{ g} = \boxed{50.0 \text{ g}}$$

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