

Stoichiometry Practice

1. Find the mass of zinc required to produce 5 moles of hydrogen gas in the following reaction:

$$Zn + 2HCl \rightarrow ZnCl_2 + H_2$$
 5 mol H_2 $\times \frac{1 \text{ mol Zn}}{1 \text{ mol } H_2} = 5 \text{ mol Zn} \times \frac{65.38 \text{ g/mol}}{1 \text{ mol}} = 326.9 \text{ g Zn}$
2. How many grams of aluminum chloride are produced when aluminum reacts with 10 grams of hydrochloric acid (HCl)?

$$2Al + 6HCl \rightarrow 2AlCl_3 + 3H_2$$

$$10 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.46 \text{ g/mol}} \times \frac{2 \text{ mol AlCl}_3}{6 \text{ mol HCl}} \times \frac{133.49 \text{ g/mol}}{1 \text{ mol}} = 12.1 \text{ g AlCl}_3$$
3. How many grams of H_2 would form in a reaction with hydrogen sulfide to produce 10 grams of H_2O ?

$$2H_2S + 3O_2 \rightarrow 2H_2O + 2SO_2$$

$$10 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g/mol}} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.016 \text{ g/mol}}{1 \text{ mol}} = 1.12 \text{ g H}_2$$
4. How many grams of carbon, will 200 g of CO_2 will react with 100 g of aluminum metal?

$$2Al + 3CO_2 \rightarrow Al_2O_3 + 3C$$

$$200 \text{ g CO}_2 \times \frac{1 \text{ mol CO}_2}{44.01 \text{ g/mol}} \times \frac{3 \text{ mol C}}{3 \text{ mol CO}_2} \times \frac{12.01 \text{ g/mol}}{1 \text{ mol}} = 133.4 \text{ g C}$$
5. How many grams of silver bromide can be precipitated by the action of 10 g of silver nitrate with 10 g of sodium bromide?

$$AgNO_3 + NaBr \rightarrow AgBr + NaNO_3$$

$$10 \text{ g AgNO}_3 \times \frac{1 \text{ mol AgNO}_3}{169.87 \text{ g/mol}} \times \frac{1 \text{ mol AgBr}}{1 \text{ mol AgNO}_3} \times \frac{235.74 \text{ g/mol}}{1 \text{ mol}} = 14.0 \text{ g AgBr}$$

$$10 \text{ g NaBr} \times \frac{1 \text{ mol NaBr}}{102.89 \text{ g/mol}} \times \frac{1 \text{ mol AgBr}}{1 \text{ mol NaBr}} \times \frac{235.74 \text{ g/mol}}{1 \text{ mol}} = 23.0 \text{ g AgBr}$$
6. How many grams of oxygen gas do you produce if 20 g of copper(II) oxide reacts with 20 g of hydrogen gas?

$$CuO + H_2 \rightarrow Cu + H_2O$$

$$20 \text{ g CuO} \times \frac{1 \text{ mol CuO}}{79.54 \text{ g/mol}} \times \frac{1 \text{ mol O}_2}{2 \text{ mol CuO}} \times \frac{32.00 \text{ g/mol}}{1 \text{ mol}} = 4.01 \text{ g O}_2$$

$$20 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2.016 \text{ g/mol}} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2} \times \frac{32.00 \text{ g/mol}}{1 \text{ mol}} = 1.59 \text{ g O}_2$$

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Stoichiometry Worksheet #1: Worked Solutions

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Answers To Stoichiometry Worksheet

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12.4: Mass-Mass Stoichiometry Last updated; Save as PDF Page ID 53792; Mass to

Mass Problems; Summary; Contributors and Attributions; I want to send 250 grams of shelled walnuts to a friend (don't ask why - just go with the question). How many walnuts in shells do I need to buy? To figure this out, I need to know how much the shell of a walnut weighs (about 40%) of the total weight of the ...

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CHM 130 Stoichiometry Worksheet The following flow chart may help you work stoichiometry problems. Remember to pay careful attention to what you are given, and what you are trying to find. 1. Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide: C 6 H 12 O 6 (s) 2 C 2 H 5 OH (l) + 2 CO 2 (g) A. Calculate the mass of ethanol produced if ...

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the limiting reagent in problem #2? d)How much of the excess reagent will be left over after the reaction is complete? Find the molarity of the ...

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Worksheet Mole Mass Problems Answers

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