

1 2 3 4 5  
 KEEP  
 CALM  
 IT'S  
 EASY
 
 Name: \_\_\_\_\_ Period: \_\_\_\_\_

Bonus points for being 1<sup>st</sup> to get all 3 tac toe and 1<sup>st</sup> to get all 9!

water    H <sub>2</sub> O(l) Grams    54 Formula mass    _____ g/mol moles    _____	Balance the following reaction $\_\_ \text{H}_2 + \_\_ \text{O}_2 \rightarrow \_\_ \text{H}_2\text{O}$ What type of reaction is this?	<table border="1" style="width: 100%; border-collapse: collapse;"> <tr> <td style="text-align: center;">2</td> <td style="text-align: center;">1</td> <td style="text-align: center;">2</td> <td style="text-align: center;">1</td> </tr> <tr> <td style="text-align: center;">Al</td> <td style="text-align: center;">+</td> <td style="text-align: center;">CuSO<sub>4</sub></td> <td style="text-align: center;">→</td> </tr> <tr> <td style="text-align: center;">2</td> <td style="text-align: center;">3</td> <td style="text-align: center;">1</td> <td style="text-align: center;">1</td> </tr> <tr> <td style="text-align: center;">Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></td> <td style="text-align: center;">+</td> <td style="text-align: center;">3</td> <td style="text-align: center;">Cu</td> </tr> <tr> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> </tr> <tr> <td style="text-align: center;">Fe</td> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> <td style="text-align: center;">+</td> </tr> <tr> <td style="text-align: center;">=</td> <td style="text-align: center;">100</td> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> </tr> </table>	2	1	2	1	Al	+	CuSO <sub>4</sub>	→	2	3	1	1	Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	+	3	Cu	+	100	+	100	Fe	+	100	+	=	100	+	100
2	1	2	1																											
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+	100	+	100																											
Fe	+	100	+																											
=	100	+	100																											
Balance the following reaction $\_\_ \text{NaCl} + \_\_ \text{F}_2 \rightarrow \_\_ \text{NaF} + \_\_ \text{Cl}_2$ What type of reaction is this?	<table border="1" style="width: 100%; border-collapse: collapse;"> <tr> <td style="text-align: center;">2</td> <td style="text-align: center;">7</td> <td style="text-align: center;">4</td> <td style="text-align: center;">6</td> </tr> <tr> <td style="text-align: center;">C<sub>2</sub>H<sub>6</sub></td> <td style="text-align: center;">+</td> <td style="text-align: center;">O<sub>2</sub></td> <td style="text-align: center;">→</td> </tr> <tr> <td style="text-align: center;">12</td> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> <td style="text-align: center;">+</td> </tr> <tr> <td style="text-align: center;">Fe</td> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> <td style="text-align: center;">+</td> </tr> <tr> <td style="text-align: center;">=</td> <td style="text-align: center;">100</td> <td style="text-align: center;">+</td> <td style="text-align: center;">100</td> </tr> </table>	2	7	4	6	C <sub>2</sub> H <sub>6</sub>	+	O <sub>2</sub>	→	12	+	100	+	Fe	+	100	+	=	100	+	100	oxygen    O <sub>2</sub> Grams    96 Formula mass    _____ g/mol moles    _____								
2	7	4	6																											
C <sub>2</sub> H <sub>6</sub>	+	O <sub>2</sub>	→																											
12	+	100	+																											
Fe	+	100	+																											
=	100	+	100																											
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2	1	2	1																											
Ag	+	Cl <sub>2</sub>	→																											
20	+	100	+																											
Fe	+	100	+																											
=	100	+	100																											

A



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